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Chapter 14

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=  $-\log[\text{H}^+]$ ; pOH

=  $-\log[\text{OH}^-]$ ; At

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25 °C, pH + pOH  
= 14.00; For

Exercise 13.37:

a.  $\text{pH} = -\log[\text{H}^+]$   
 $= -\log(1.0 \times 10^{-7}) = 7.00;$

$\text{pOH} = 14.00 - \text{pH}$   
 $= 14.00 - 7.00 =$   
 $7.00$

b.  $\text{pH} = -\log(8.3 \times 10^{-16}) = 15.08;$

$\text{pOH} = 14.00 - 15.08 = -1.08$

c.  $\text{pH} = -\log(12) =$

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$$\begin{aligned} & -1.08; \text{pOH} = \\ & 14.08 - (-1.08) \\ & = 15.08 \text{ d.} \end{aligned}$$

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Och1g9 7d7ryh.  
... The result

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of this is the one-significant-figure answer of 0.5. Next, the multiplication/division rule is applied to  $0.5/0.50$ . ...

Thus  $1 \text{ ha} = 2.5 \text{ acre}$ . From the previous question, we are given that  $1 \text{ grain troy} =$

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0.0648 g = 1  
grain ap. So the  
two are the  
same.The ...

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Chapter 4

Answers -

Zumdahl a.  $\text{KMnO}_4$

is composed of

$\text{K}^+$  and  $\text{MnO}_4^-$

ions. Assign

oxygen a value

of  $-2$ , which

gives manganese

a  $+7$  oxidation

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state since the sum of oxidation states for all atoms in  $\text{MnO}_4^-$  must equal the -1 charge on  $\text{MnO}_4^-$ .  
K, +1; O, -2;  
Mn, +7.

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